Volunteer Monitoring Quarterly Data Report

Beaver Lake Watershed - Benton, Washington and Madison counties in Arkansas

May 2018

Monitoring Period: May 6-20, 2018

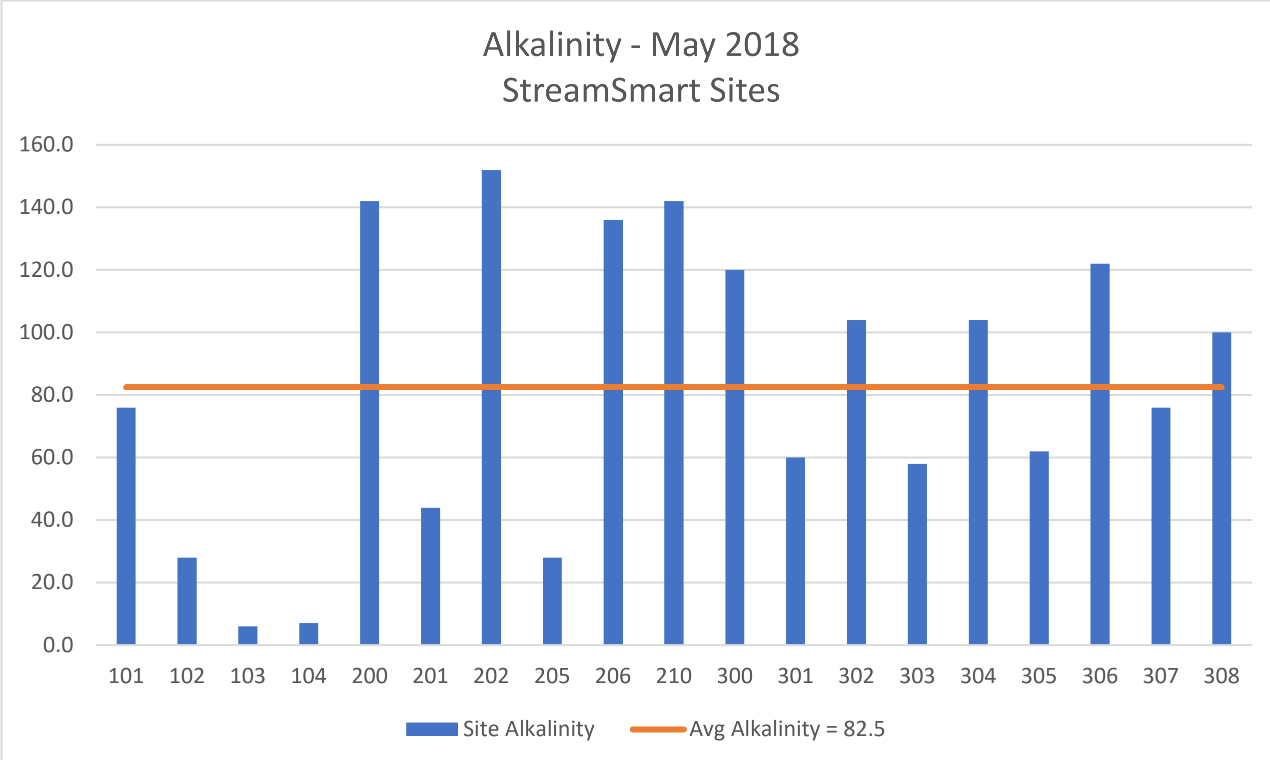
A project of Ozarks Water Watch in Arkansas

Prepared by: Angela Danovi
StreamSmart Coordinator
Ozarks Water Watch
1200 W. Walnut Street, Suite 3405
Rogers, AR 72745

StreamSmart Monitoring Sites – 2018

Site Number	Site Name	Lat/Long	Watershed
101	West Fork (Baptist Ford Bridge)	35.982714, -94.173129	West Fork
102	West Fork (Brentwood Park)	35.865723, -94.117257	West Fork
103	Baldwin Creek Near St. Paul	35.822256, -93.758937	Headwaters
104	White River Near St. Paul	35.818676, -93.779774	Headwaters
107	War Eagle Creek	35.888319, -93.679017	War Eagle
108	Ogden Creek	35.887777, -93.679069	War Eagle
200	Ward Slough	35.997178, -94.173949	West Fork
201	Middle Fork of W.R. at Harris Rd	35.995825, -94.072894	Middle Fork
202	College Branch - U of A		West Fork
205	Hock Creek	36.022453, -93.859784	Richland Creek
206	Spout Spring Branch	36.055019, -94.161107	West Fork
210	Town Branch (White River Ball fields)	36.043179, -94.135852	West Fork
300	Brush Creek	36.131947, -93.947956	Beaver Reservoir
301	War Eagle Creek (Huntsville)	36.149997, -93.740137	War Eagle
302	Glade Creek	36.159851, -93.811690	War Eagle
303	Clear Creek	36.195153, -93.789276	War Eagle
304	Clifty Creek	36.239342, -93.907653	War Eagle
305	War Eagle (Mill)	36.267597, -93.943130	War Eagle
306	Prairie Creek	36.341208, -94.096513	Beaver Reservoir
307	Holman Creek Upstream of Huntsville	36.104418, -93.756750	War Eagle
308	Holman Creek Downstream of Huntsville	36.124453, -93.734211	War Eagle

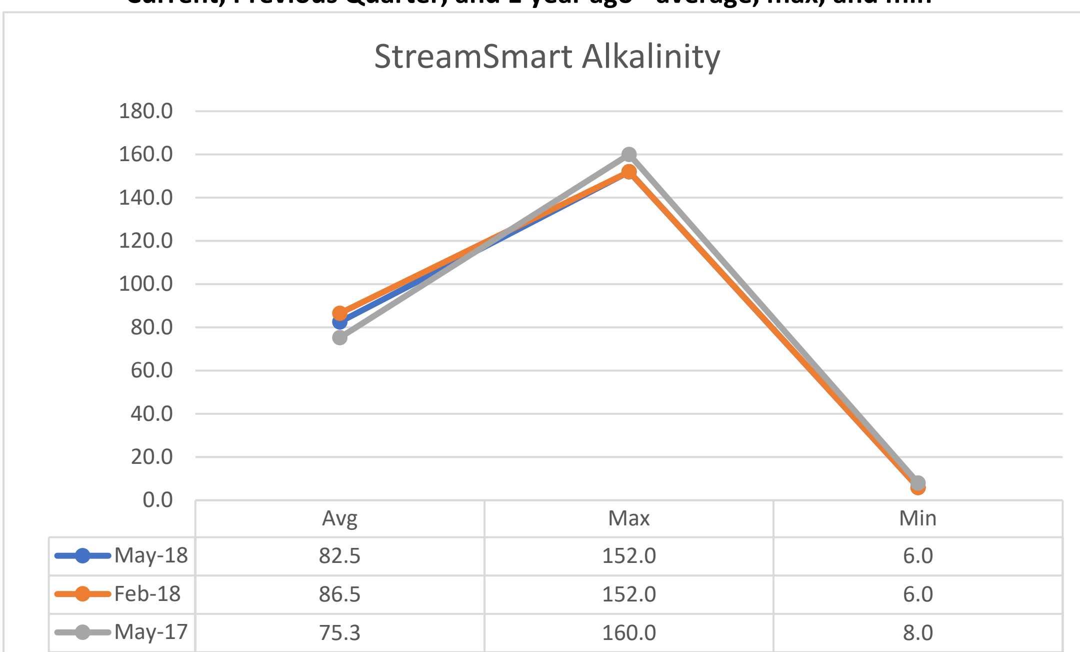
Alkalinity Data – May 2018 StreamSmart



May 2018 Data

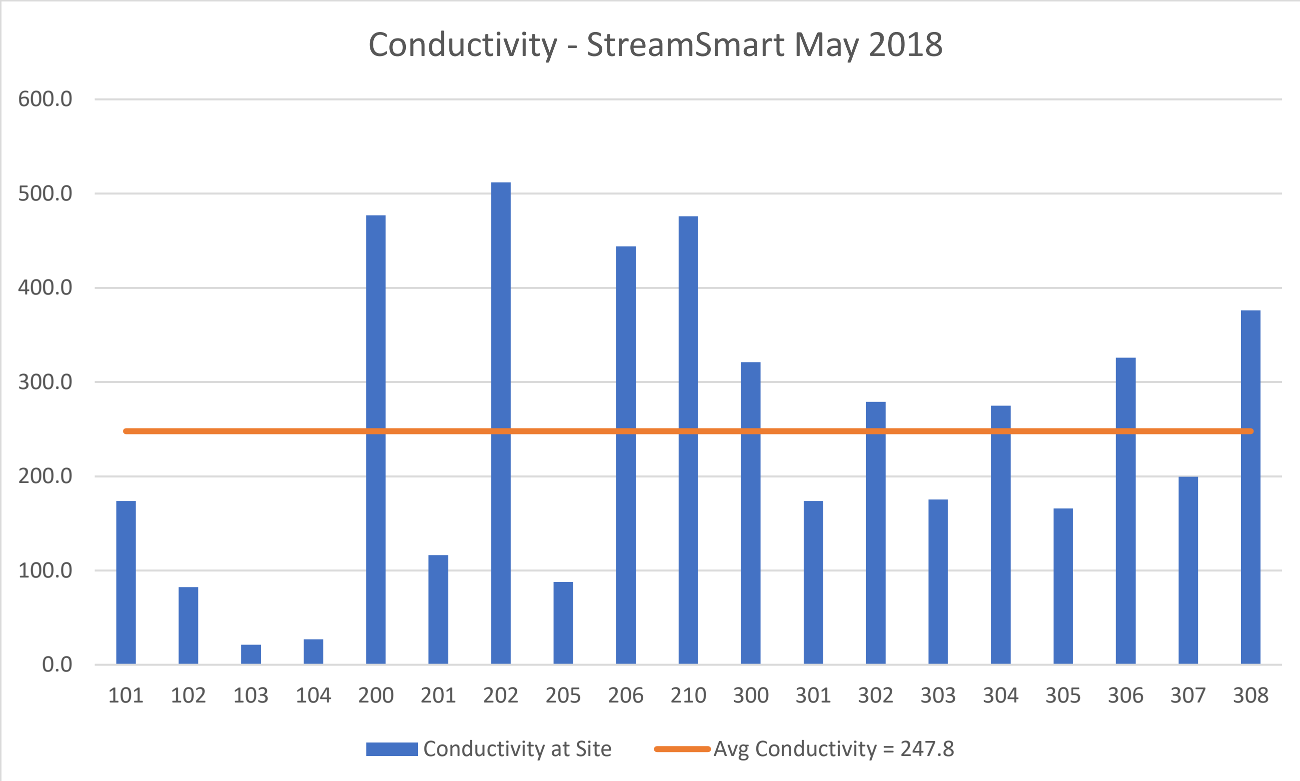
Site	Alkalinity
101	76.0
102	28.0
103	6.0
104	7.0
200	142.0
201	44.0
202	152.0
205	28.0
206	136.0
210	142.0
300	120.0
301	60.0
302	104.0
303	58.0
304	104.0
305	62.0
306	122.0
307	76.0
308	100.0

Current, Previous Quarter, and 1 year ago - average, max, and min



	Max	Min
May-18	202 - Mullins Branch	103 - Baldwin Creek
Feb-18	306 - Prairie Creek	103 - Baldwin Crek
May-17	206 - Spout Spring Branch	104 - White River Near St. Paul

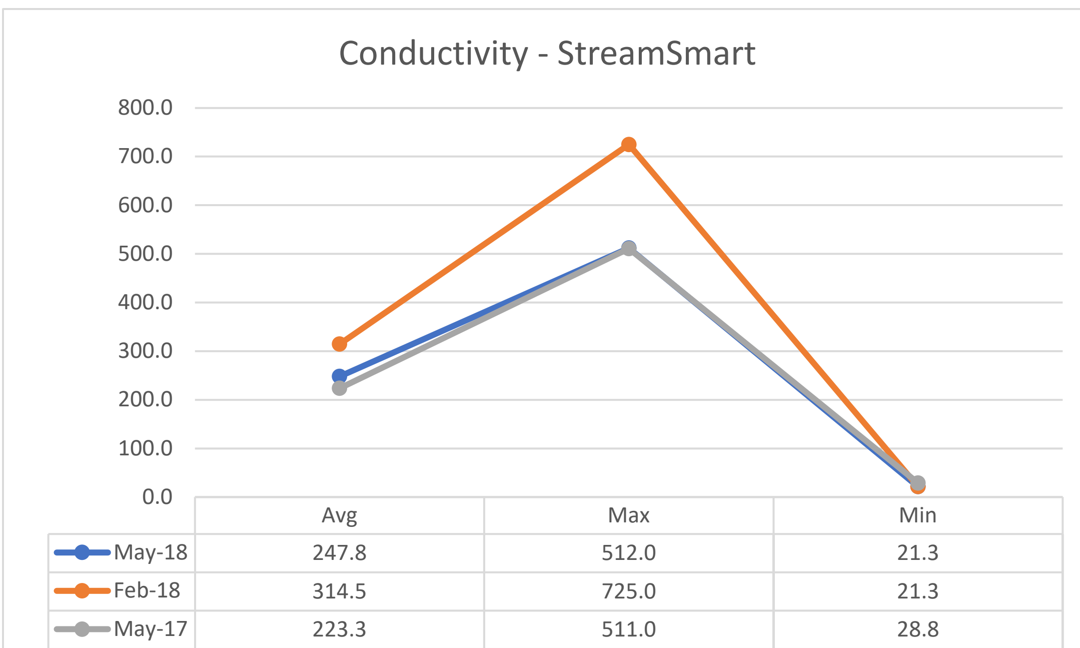
Conductivity Data – May 2018 StreamSmart



May 2018 Data

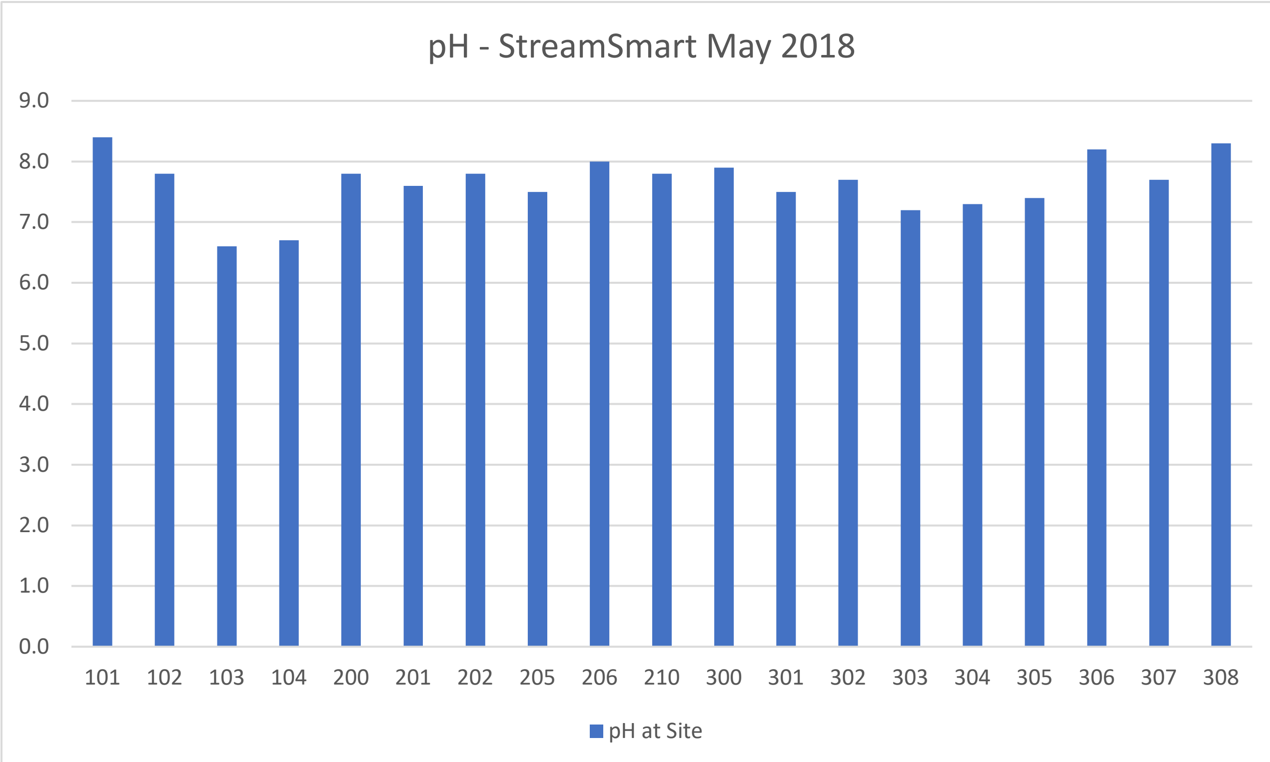
Site	Conductivity
101	173.7
102	82.5
103	21.3
104	26.9
200	477.0
201	116.3
202	512.0
205	87.8
206	444.0
210	476.0
300	321.0
301	173.6
302	279.0
303	175.4
304	275.0
305	165.9
306	326.0
307	199.5
308	376.0

Current, Previous Quarter, and 1 year ago - average, max, and min



	Max	Min
Feb-18	May-18	202 - Mullins Creek
Nov-17	Feb-18	202 - Mullins Creek
Feb-17	May-17	206 - Spout Spring Branch

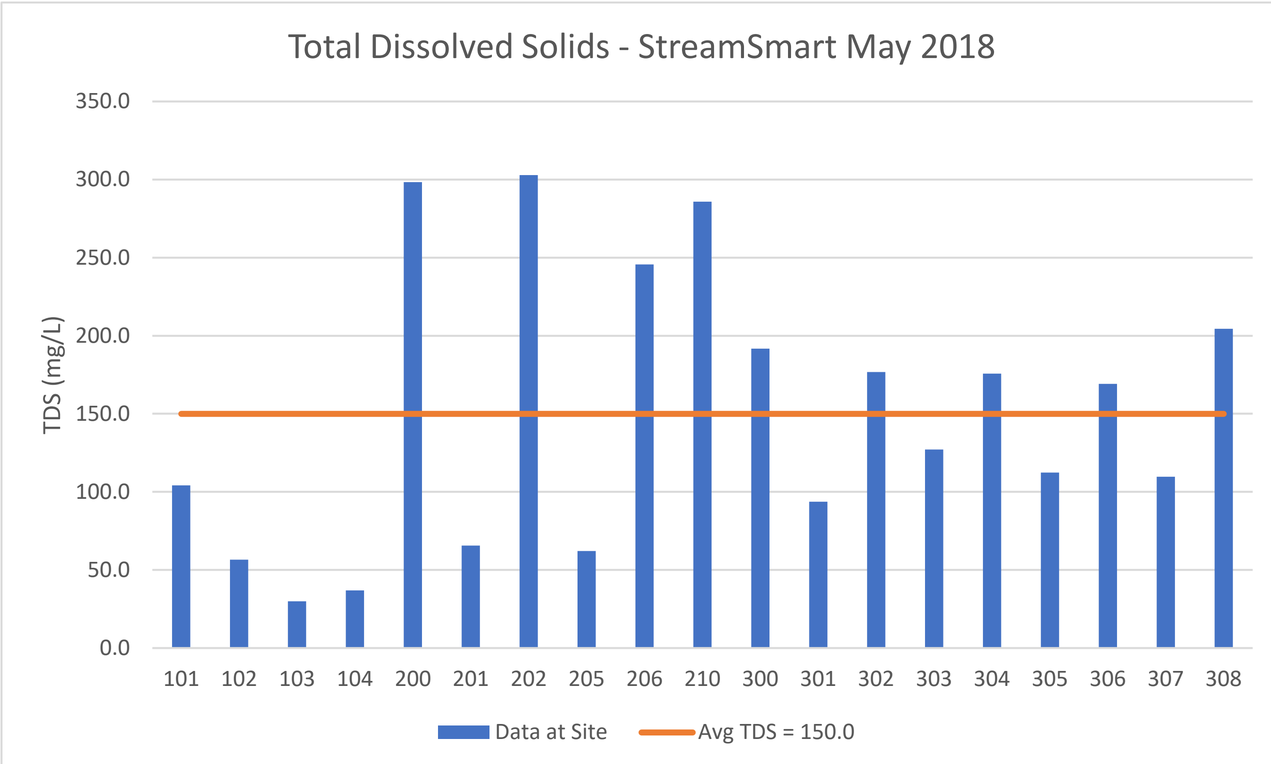
pH Data – May 2018 StreamSmart



May 2018 Data

Site	pH
101	8.4
102	7.8
103	6.6
104	6.7
200	7.8
201	7.6
202	7.8
205	7.5
206	8.0
210	7.8
300	7.9
301	7.5
302	7.7
303	7.2
304	7.3
305	7.4
306	8.2
307	7.7
308	8.3

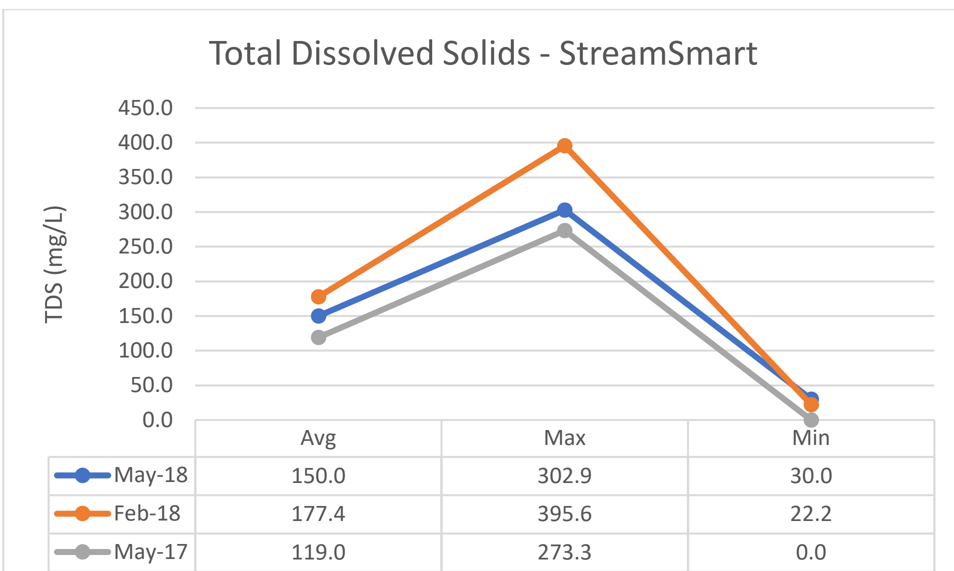
Total Dissolved Solids – May 2018 StreamSmart



May 2018 Data

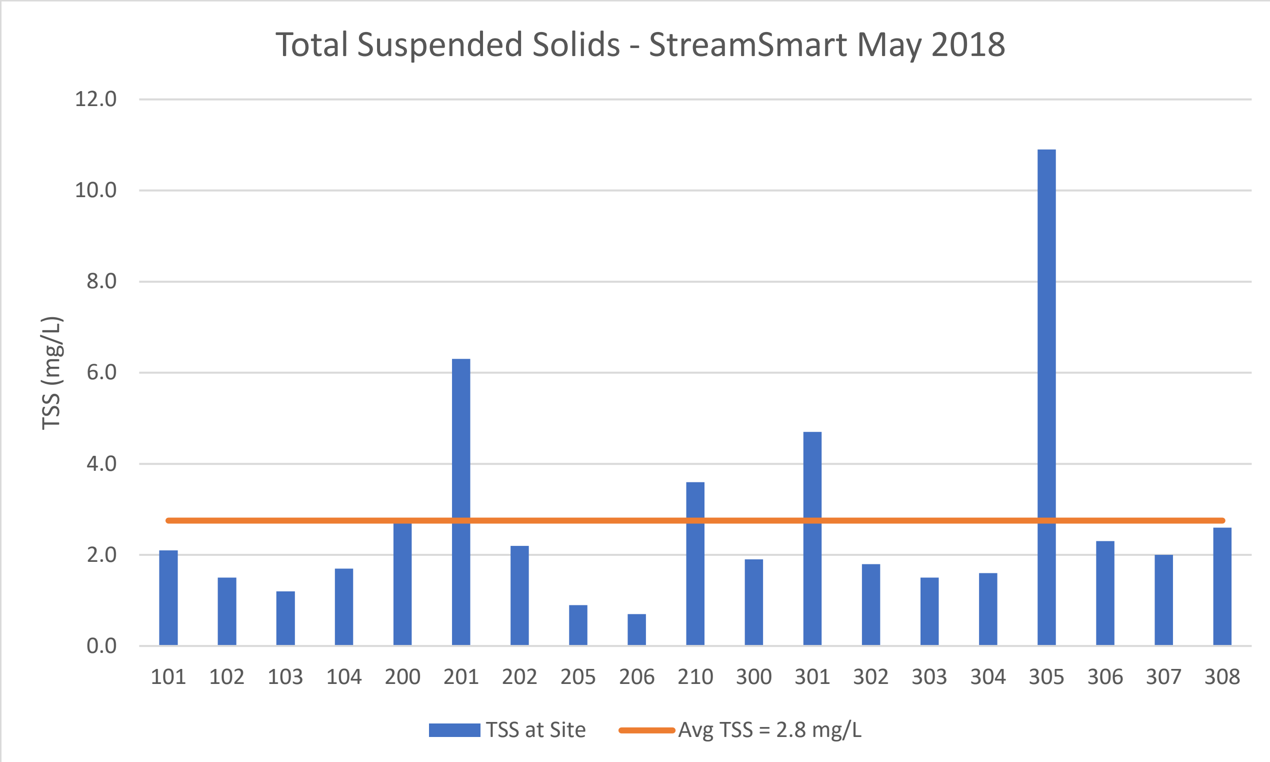
Site	TDS (mg/L)
101	104.2
102	56.7
103	30.0
104	36.9
200	298.4
201	65.6
202	302.9
205	62.2
206	245.6
210	285.8
300	191.8
301	93.8
302	176.7
303	127.1
304	175.8
305	112.4
306	169.1
307	109.8
308	204.4

Current, Previous Quarter, and 1 year ago - average, max, and min



	Max	Min
May-18	202 - Mullins Branch	103 - Baldwin Creek
Feb-18	202 - Mullins Creek	104 - White River near St. Paul
May-17	206 - Spout Spring Branch	104 - White River near St. Paul

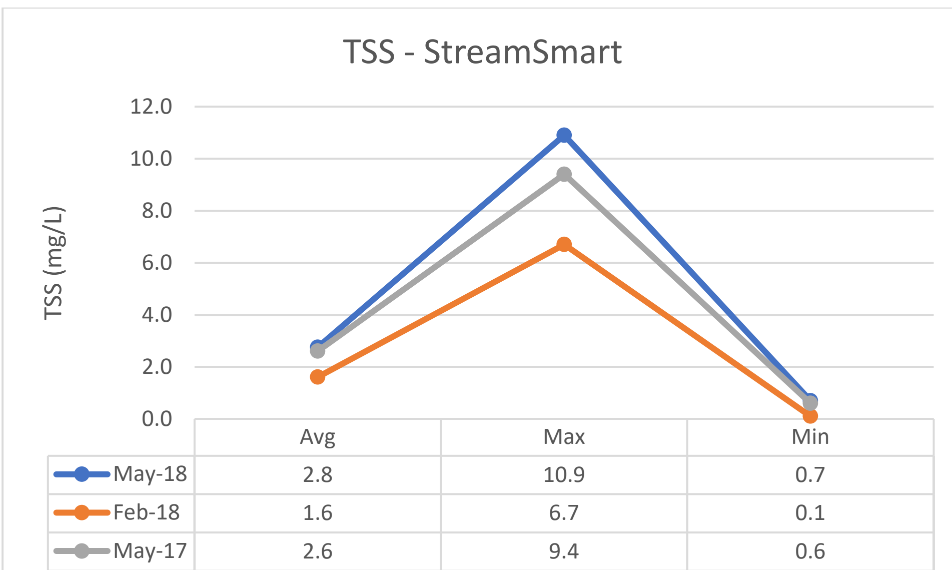
Total Suspended Solids – February 2018 StreamSmart



May 2018 Data

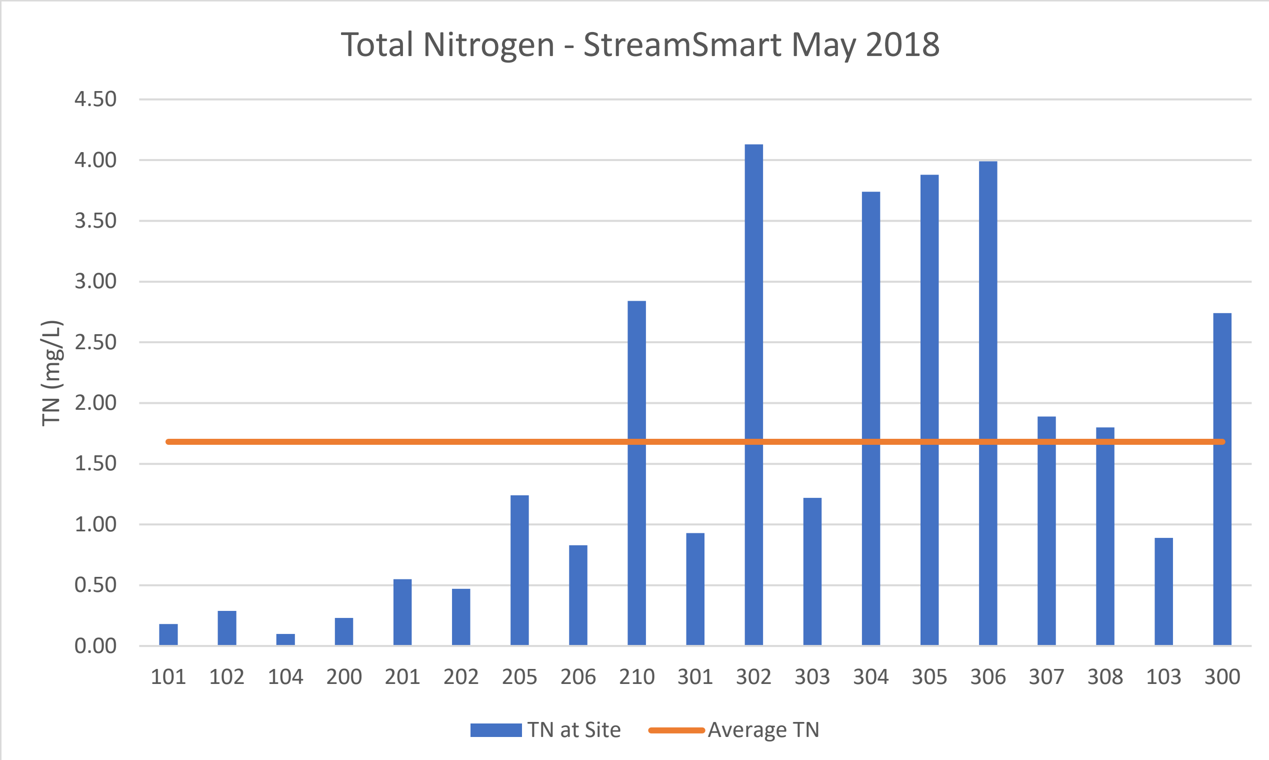
Site	TSS (mg/L)
101	2.1
102	1.5
103	1.2
104	1.7
200	2.8
201	6.3
202	2.2
205	0.9
206	0.7
210	3.6
300	1.9
301	4.7
302	1.8
303	1.5
304	1.6
305	10.9
306	2.3
307	2.0
308	2.6

Current, Previous Quarter, and 1 year ago - average, max, and min



	Max	Min
May-18	305 - War Eagle Mill	206 - Spout Spring Branch
Feb-18	307 - Holman Creek Upstream of Huntsville	210 - Town Branch
May-17	305 - War Eagle Mill	300 - Brush Creek

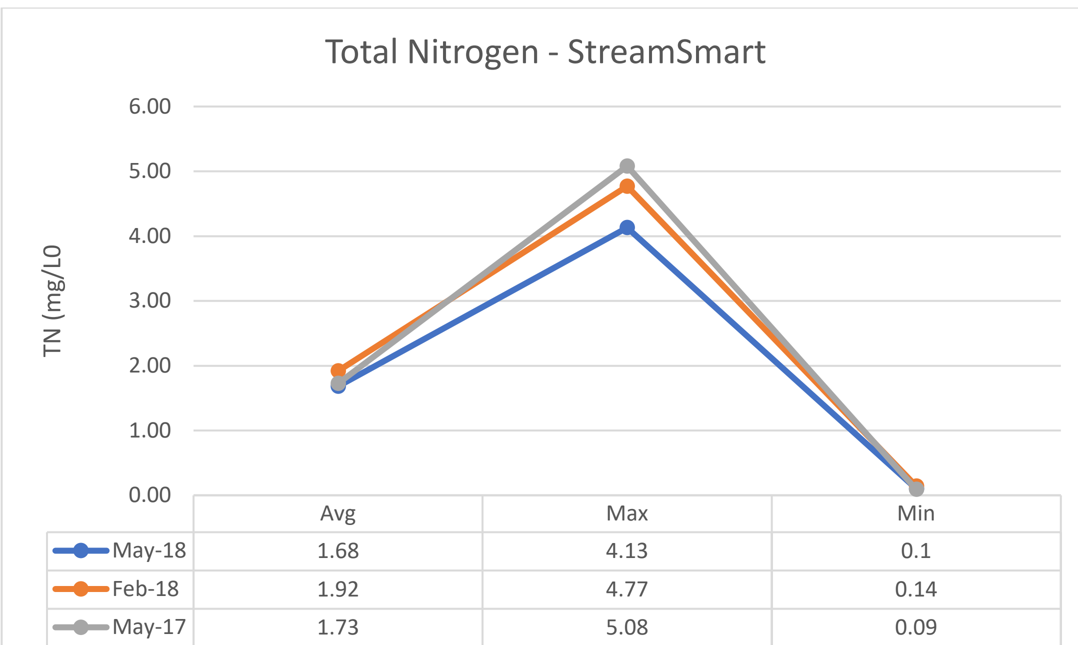
Total Nitrogen – May 2018 StreamSmart



May 2018 Data

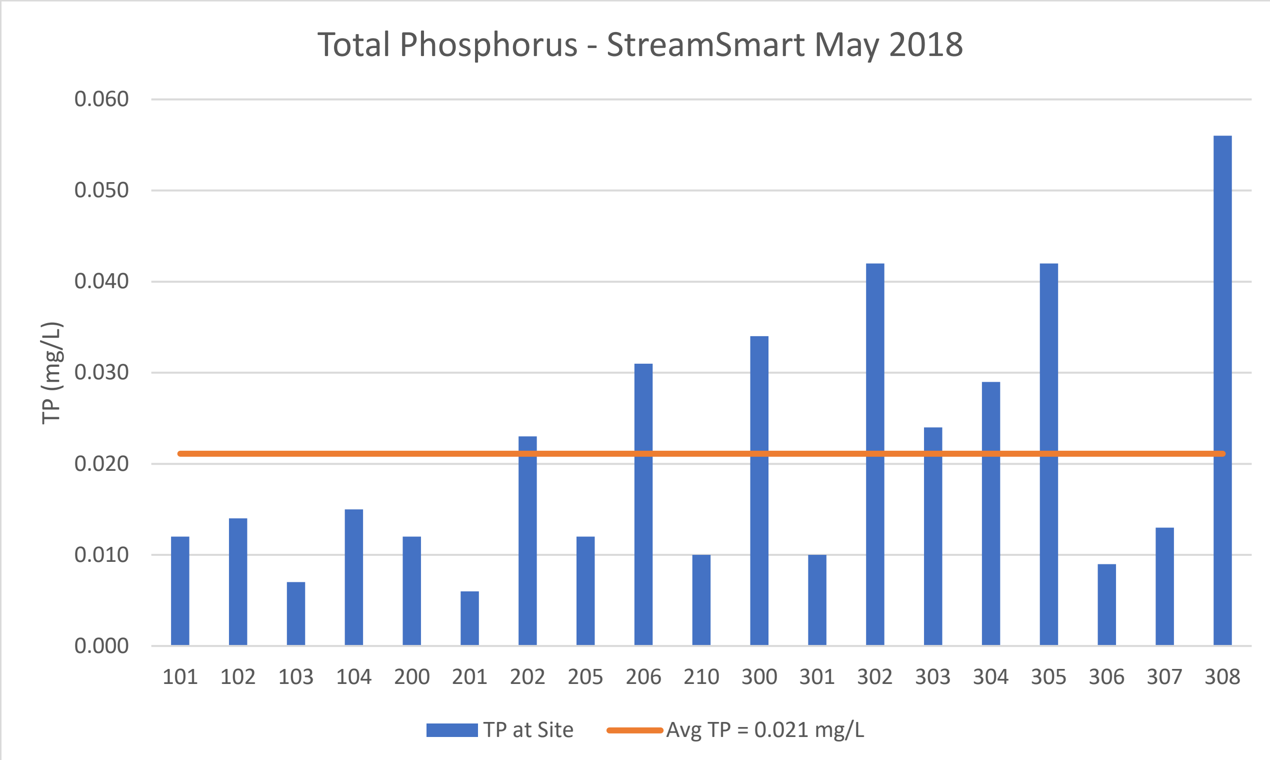
Site	TN (mg/L)
101	0.18
102	0.29
104	0.10
200	0.23
201	0.55
202	0.47
205	1.24
206	0.83
210	2.84
301	0.93
302	4.13
303	1.22
304	3.74
305	3.88
306	3.99
307	1.89
308	1.80
103	0.89
300	2.74

Current, Previous Quarter, and 1 year ago - average, max, and min



	Max	Min
May-18	302 - Glade Creek	104 - White River near St. Paul
Feb-18	308 - Holman Creek Downstream of Huntsville	103 - Baldwin Creek
May-17	300 - Brush Creek	103 - Baldwin Creek

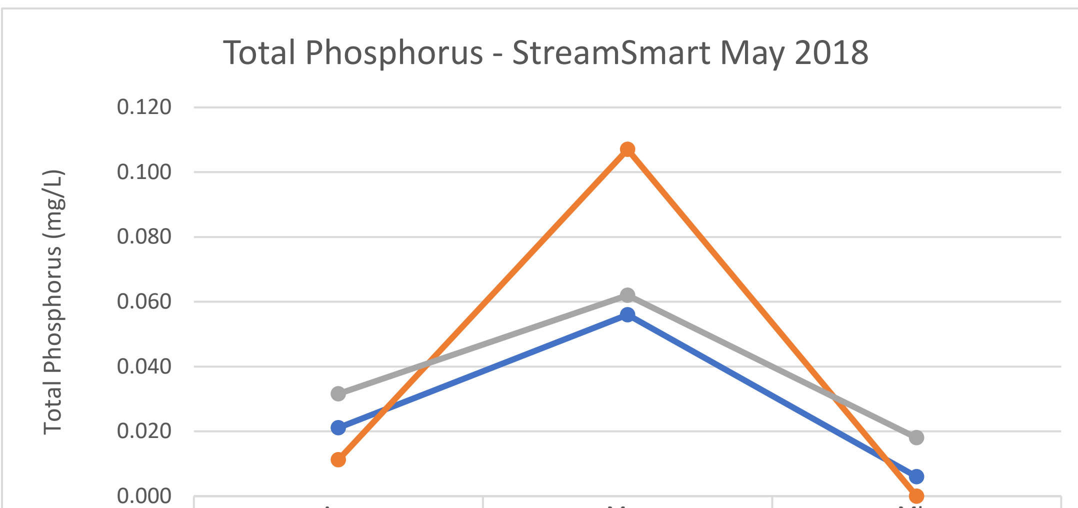
Total Phosphorus – May 2018 StreamSmart



May 2018 Data

Site	TP (mg/L)
101	0.012
102	0.014
103	0.007
104	0.015
200	0.012
201	0.006
202	0.023
205	0.012
206	0.031
210	0.010
300	0.034
301	0.010
302	0.042
303	0.024
304	0.029
305	0.042
306	0.009
307	0.013
308	0.056

Current, Previous Quarter, and 1 year ago - average, max, and min



	Avg	Max	Min
May-18	0.021	0.056	0.006
Feb-18	0.011	0.107	0.000
May-17	0.032	0.062	0.018

	Max Site Number - Location	Min Site Number - Location
May-18	308 - Holman Creek DS of Huntsvillie	201 - MF of WR at Harris Rd.
Feb-18	308 - Holman Creek DS of Huntsvillie	101 - WF at Baptist Ford Bridge/102 - WF at Brentwood Park
May-17	202 - Mullins Creek	103 - Baldwin Creek / Site 104 - W.R. Near St. Paul

What is Alkalinity?

Alkalinity is the water's capacity to resist changes in pH that would make the water more acidic. Alkalinity refers to the capability of water to neutralize acid. The alkalinity of natural water is determined by the soil and bedrock through which it passes. The main sources for natural alkalinity are rocks which contain carbonate, bicarbonate, and hydroxide compounds. Limestone is rich in carbonates, so waters flowing through limestone regions or bedrock containing carbonates generally have high alkalinity - hence good buffering capacity. Conversely, areas rich in granites and some conglomerates and sandstones may have low alkalinity and, therefore, poor buffering capacity.

Alkalinity is important for fish and aquatic life because it protects or buffers against rapid pH changes. Living organisms, especially aquatic life, function best in a pH range of 6.0 to 9.0. Alkalinity is a measure of how much acid can be added to a liquid without causing a large change in pH. Higher alkalinity levels in surface waters will buffer acid rain and other acid wastes and prevent pH changes that are harmful to aquatic life.

http://www.water-research.net/index.php/the-role-of-alkalinity-citizen-monitoring

What is Conductivity:

Conductivity is a measurement of the ability of an aqueous solution to carry an electrical current. Typically, the units of measure are microhms/cm (uohms/cm) or microsiemens/cm (uS/cm). Conductivity or specific conductance is a measure of the ability of a fluid to carry a charge which is directly related to the concentration of dissolved substances. As the total dissolved substances in the water increases, the conductivity of the water also increases.

http://www.water-research.net/index.php/drinking-water-testing-and-conductivity-of-water

What is pH:

pH measurements run on a scale from 0 to 14, with 7.0 considered neutral. Solutions with a pH below 7.0 are considered acids. Solutions with a pH above 7.0, up to 14.0 are considered bases. Living organisms, especially aquatic life, function best in a pH range of 6.0 to 9.0.

The pH scale is logarithmic, so every one-unit change in pH represents a ten-fold change in acidity. In other words, pH 6.0 is ten times more acidic than pH 7.0; pH 5 is one hundred times mores acidic than pH 7.0.

The pH of a body of water is affected by several factors including the bedrock and soil composition through which the water moves. Some rock types such as limestone can, to an extent, neutralize acid. Another factor which affects the pH is the amount of plant growth and organic material within a body of water. When this material decomposes carbon dioxide is released. The carbon dioxide combines with water to form carbonic acid. Although this is a weak acid, large amounts of it will lower the pH. A third factor which determines the pH of a body of water is the dumping of chemicals into the water by individuals, industries, and communities.

Changes in the pH value of water are important to many organisms. Most organisms have adapted to life in water of a specific pH and may die if it changes even slightly. This is especially true of aquatic macroinvertebrates and fish eggs and fry.

http://www.water-research.net/index.php/ph-in-the-environment

What is Total Dissolved Solids?

A water quality parameter defining the concentration of dissolved organic and inorganic chemicals in water. After suspended solids are filtered from water and water is evaporated, dissolved solids are the remaining residue. An elevated total dissolved solids concentration does not mean that the water is a health hazard, but it does mean the water may have aesthetic problems, such as taste and odor, or cause nuisance problems.

http://www.water-research.net/index.php/water-treatment/tools/total-dissolved-solids

What is Total Suspended Solids?

The suspended or colloidal particles, commonly referred to as total suspended solids (TSS), are all the extremely small suspended solids in water which will not settle out by gravity. TSS is measured on a sample of water (which has been settled) and are those particles which will not pass through a very fine filter.

TSS in streams in northwest Arkansas usually range from 0.1 to 20 mg/L but can get as high as 500 mg/L during storm flows because the faster water moves the more sediment it can carry and the more force it has to cause erosion of the stream banks and channel.

What is Total Phosphorus?

Phosphorus occurs naturally in rocks and other mineral deposits. During the natural process of weathering, the rocks gradually release the phosphorus as phosphate ions which are soluble in water and the mineralize phosphate compounds breakdown. Phosphorus is one of the key elements necessary for the growth of plants and animals and in lake ecosystems it tends to be the growth-limiting nutrient.

Total phosphorus is a measure of all the forms of phosphorus in the sample (orthophosphate, condensed phosphate, and organic phosphate). This is accomplished by first "digesting" (heating and acidifying) the sample to convert all the other forms to orthophosphate. Then the orthophosphate is measured by the ascorbic acid method. Because the sample is not filtered, the procedure measures both dissolved and suspended orthophosphate. Monitoring phosphorus is challenging because it involves measuring very low concentrations down to 0.01 milligram per liter (mg/L) or even lower. Even such very low concentrations of phosphorus can have a dramatic impact on streams.

What is Total Nitrogen?

There are three forms of nitrogen that are commonly measured in water bodies: ammonia, nitrates and nitrites. Total nitrogen is the sum of total kjeldahl nitrogen (ammonia, organic and reduced nitrogen) and nitrate-nitrite. It can be derived by monitoring for organic nitrogen compounds, free-ammonia, and nitrate-nitrite individually and adding the components together. An acceptable range of total nitrogen is 2 mg/L to 6 mg/L, though variations from this range can occur. We measure total Nitrogen as part of our on-going monitoring of nutrients concentrations in surface water.